

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide and hydronium ion concentrations, pH, and pOH, and vice versa, for strong acids and bases.
3. Calculate the pH of weak acids.

Vocabulary

- | | | |
|-------------------|-----------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. strong acid/base |
| 4. hydronium ion | 5. weak acid/base | 6. pH |

Key Equations

- | | |
|---|---|
| 1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (page 714) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (page 714) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (page 714) | 4. $\text{pH} + \text{pOH} = 14.00$ (page 714) |

Part A: Completion

Use the appropriate number of significant figures in your answers. If the answer requires that you round to the nearest integer, indicate whether you are rounding up or down.

1. What products are **1** **2** formed by the autoionization of water? Write the chemical formula of the products of these acids produced in **3** **4** water.
2. Hydrochloric acid has a $\text{pH} = 2.35$. What is the concentration of hydroxide ions in this acid? Write **5** **6** **7** **8** **9** **10** **11** **12** **13** **14** **15** **16** **17** **18** **19** **20** **21** **22** **23** **24** **25** **26** **27** **28** **29** **30** **31** **32** **33** **34** **35** **36** **37** **38** **39** **40** **41** **42** **43** **44** **45** **46** **47** **48** **49** **50** **51** **52** **53** **54** **55** **56** **57** **58** **59** **60** **61** **62** **63** **64** **65** **66** **67** **68** **69** **70** **71** **72** **73** **74** **75** **76** **77** **78** **79** **80** **81** **82** **83** **84** **85** **86** **87** **88** **89** **90** **91** **92** **93** **94** **95** **96** **97** **98** **99** **100**

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