

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities (correct)?
 (A) pressure, concentration (B) pressure, pH
 (C) molar fraction (D) dimensional analysis
2. Which of the following is CORRECT? (A) concentration of ions increases as temperature increases

(A) $K_w = 10^{-14}$ (B) $K_w = 10^{-16}$

(C) $10^{-14} \text{ mol/L} < [\text{H}^+] < 10^{-7} \text{ mol/L}$ (D) $10^{-14} \text{ mol/L} < [\text{OH}^-] < 10^{-7} \text{ mol/L}$
 (E) $10^{-14} \text{ mol/L} < [\text{H}^+] < 10^{-7} \text{ mol/L}$ (F) none of the above

3. In a given solution, the following equilibrium is established: $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$. If the concentration of H_3O^+ is 10^{-7} mol/L , what is the concentration of OH^- ?

(A) 10^{-7} mol/L (B) 10^{-14} mol/L (C) 10^{-16} mol/L (D) 10^{-17} mol/L (E) 10^{-18} mol/L

(F) 10^{-19} mol/L (G) 10^{-20} mol/L (H) 10^{-21} mol/L (I) 10^{-22} mol/L

4. In an acidic solution, the concentration of H^+ is 10^{-3} mol/L . What is the concentration of OH^- ?

(A) 10^{-11} mol/L (B) 10^{-12} mol/L (C) 10^{-13} mol/L (D) 10^{-14} mol/L

5. Which of the following is CORRECT? (A) $K_w = 10^{-14}$ (B) $K_w = 10^{-16}$

(C) $10^{-14} \text{ mol/L} < [\text{H}^+] < 10^{-7} \text{ mol/L}$ (D) $10^{-14} \text{ mol/L} < [\text{OH}^-] < 10^{-7} \text{ mol/L}$

(E) $10^{-14} \text{ mol/L} < [\text{H}^+] < 10^{-7} \text{ mol/L}$ (F) $10^{-14} \text{ mol/L} < [\text{OH}^-] < 10^{-7} \text{ mol/L}$

6. A solution is made by mixing 100 mL of 0.1 M HCl with 100 mL of 0.1 M NaOH . What is the pH of the resulting solution?

(A) 1 (B) 7 (C) 13 (D) 14 (E) 15 (F) 16 (G) 17 (H) 18 (I) 19

(J) 10 (K) 11 (L) 12 (M) 13 (N) 14 (O) 15 (P) 16 (Q) 17 (R) 18 (S) 19

7. In a weak acid solution, the concentration of H^+ is 10^{-3} mol/L . What is the concentration of OH^- ?

(A) 10^{-11} mol/L (B) 10^{-12} mol/L (C) 10^{-13} mol/L (D) 10^{-14} mol/L

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