



Wisconsin Department of Natural Resources
Wastewater Operator Certification

Basic General Wastewater Study Guide
August 2015 Edition (Revised June 2016)



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Operator Certification Program ☐
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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
 18g H₂O / mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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 what is the limiting reagent if 8.52 g are actually formed what is the

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 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$
 if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g O}_2}{32 \text{ g/mole}} = 0.666 \text{ moles O}_2$
 $\frac{21.3 \text{ g O}_2}{16 \text{ g/mole}} = 1.33 \text{ moles O}$

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web percent yield is a measure of the efficiency of the experimental design yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 4 Al + 3 O₂ → 2 Al₂O₃
 b determine the theoretical yield of Al₂O₃
 $\frac{160 \text{ g Al}}{27 \text{ g/mol Al}} = 5.93 \text{ mol Al}$
 $\frac{101.964 \text{ g Al}_2\text{O}_3}{302.3 \text{ g/mol Al}_2\text{O}_3} = 0.337 \text{ mol Al}_2\text{O}_3$
 $\frac{26.982 \text{ g Al}}{27 \text{ g/mol Al}} = 1 \text{ mol Al}$
 c determine the percent yield
 $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{260.0 \text{ g}} \times 100 = 100\%$
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$ theoretical = 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield = 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H₂ are left over 13

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