

	Reaction (Oxidised form + ne ⁻)	→	Reduced form)	E ^o / V
↑ Increasing strength of oxidising agent	F ₂ (g) + 2e ⁻	→	2F ⁻	2.87
	Co ³⁺ + e ⁻	→	Co ²⁺	1.81
	H ₂ O ₂ + 2H ⁺ + 2e ⁻	→	2H ₂ O	1.78
	MnO ₄ ⁻ + 8H ⁺ + 5e ⁻	→	Mn ²⁺ + 4H ₂ O	1.51
	Au ³⁺ + 3e ⁻	→	Au(s)	1.40
	Cl ₂ (g) + 2e ⁻	→	2Cl ⁻	1.36
	Cr ₂ O ₇ ²⁻ + 14H ⁺ + 6e ⁻	→	2Cr ³⁺ + 7H ₂ O	1.33
	O ₂ (g) + 4H ⁺ + 4e ⁻	→	2H ₂ O	1.23
	MnO ₂ (s) + 4H ⁺ + 2e ⁻	→	Mn ²⁺ + 2H ₂ O	1.23
	Br ₂ + 2e ⁻	→	2Br ⁻	1.09
	NO ₃ ⁻ + 4H ⁺ + 3e ⁻	→	NO(g) + 2H ₂ O	0.97
	2Hg ²⁺ + 2e ⁻	→	Hg ₂ ²⁺	0.92
	Ag ⁺ + e ⁻	→	Ag(s)	0.80
	Fe ³⁺ + e ⁻	→	Fe ²⁺	0.77
	O ₂ (g) + 2H ⁺ + 2e ⁻	→	H ₂ O ₂	0.68
	I ₂ + 2e ⁻	→	2I ⁻	0.54
	Cu ⁺ + e ⁻	→	Cu(s)	0.52
	Cu ²⁺ + 2e ⁻	→	Cu(s)	0.34
	AgCl(s) + e ⁻	→	Ag(s) + Cl ⁻	0.22
	AgBr(s) + e ⁻	→	Ag(s) + Br ⁻	0.10
	2H ⁺ + 2e ⁻	→	H ₂ (g)	0.00
	Pb ²⁺ + 2e ⁻	→	Pb(s)	-0.13
	Sn ²⁺ + 2e ⁻	→	Sn(s)	-0.14
	Ni ²⁺ + 2e ⁻	→	Ni(s)	-0.25
	Fe ²⁺ + 2e ⁻	→	Fe(s)	-0.44
	Cr ³⁺ + 3e ⁻	→	Cr(s)	-0.74
Zn ²⁺ + 2e ⁻	→	Zn(s)	-0.76	
2H ₂ O + 2e ⁻	→	H ₂ (g) + 2OH ⁻ (aq)	-0.83	
Al ³⁺ + 3e ⁻	→	Al(s)	-1.66	
Mg ²⁺ + 2e ⁻	→	Mg(s)	-2.36	
Na ⁺ + e ⁻	→	Na(s)	-2.71	
Ca ²⁺ + 2e ⁻	→	Ca(s)	-2.87	
K ⁺ + e ⁻	→	K(s)	-2.93	
Li ⁺ + e ⁻	→	Li(s)	-3.05	

↓
Increasing strength of reducing agent

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