

Teacher's Edition



# Consumer Mathematics



# Novel Stars Answer Key For Consumer Math

**Kathleen Armour**



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with 1200 ml of 2M  $\text{HNO}_3$  if the final density of solution is same as pure water then molarity of the iron in resulting solution which decides the nature of the above solution is 1) 0.25M 2) 0.5M 3) 0.888M 4) none of

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$$m = \frac{M \cdot 1000}{1000d}$$

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web 1) molarity  $M = \frac{m}{V}$  mass molar mass  $\times 0.4500 \text{ L} = 0.825 \text{ g} \times 141.9579 \text{ g/mol} = 0.1172 \text{ mol}$  2) molality  $m = \frac{m}{M}$   $0.825 \text{ g} \times 141.9579 \text{ g/mol} = 0.00581158 \text{ mol}$  3) mole fraction  $n_2 = 0.825 \text{ g} \times 141.9579 \text{ g/mol} = 0.1172 \text{ mol}$   $n_1 = 0.4500 \text{ kg} = 450 \text{ g}$   $n_1 = \frac{450 \text{ g}}{18.015 \text{ g/mol}} = 24.97918401 \text{ mol}$

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solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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$$M_{mix} = \frac{m_1 \cdot V_1 + m_2 \cdot V_2 + \dots + m_n \cdot V_n}{V_{total}}$$
 1.5 x 480 + 1.2 x 520 = 480 + 520 = 1000 ml  
 1.344 M with increase in temperature which of these changes a molality b fraction of solute present in water

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