

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1V_1=P_2V_2$
Isolate for P₂: $P_2=(P_1V_1)/V_2$
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1=34.75 + 273=307.75 \text{ K}$
 $T_2=78.41 + 273=351.41 \text{ K}$
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
Get rid of P
 $V_1/T_1=V_2/T_2$
Isolate V₂: $V_2=(V_1/T_1)T_2$
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV=nRT$ will work. (n= moles and R= 0.08206 L atm/K mol)
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$ OOPS we need to convert from grams to moles!, $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$
n: Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$
Isolate V: $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

2nd Semester Final Review Guide Chemistry

R Barnett



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with no breach opening as below

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web unlined chimney note polated and i number of zones into which chimney is divided 175 2 70 4 50 0 1 pwv table 5
coefficient of dynamic 5 6 while determining the wind load on the chimney for steel chimneys clause 8 5 the value of
coefficient v shall be taken from table 7 depending upon the parameter e_l as given in 8 8

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typical chimney to be located at mumbai is taken for the example the chimney is first analysed for static wind load seismic
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top d_1 in cm 132 chimney dia bottom b_d in cm 220 chimney height total h in cm 4600 corrosion allowance mm 3 allowable
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