

Gas kinetics

molecules arrival rate R at a surface
(unit area, time)

if M_r = relative Molar mass $m = M_r \cdot \text{amu}$

$$R \propto \frac{p}{\sqrt{2mk_B T}}$$

$$m_{Ne} = 20 \cdot 1.67 \times 10^{-27} \text{ g}$$

$$R \propto \frac{p}{\sqrt{2mk_B T}} \propto \frac{1}{\sqrt{2M_r}} \frac{1}{\sqrt{1.67 \times 10^{-27} k_B}} \frac{p}{\sqrt{M_r T}}$$

$$0.399 \frac{1}{\sqrt{1.67 \times 10^{-27} \times 1.38 \times 10^{-23}}} \frac{p}{\sqrt{M_r T}}$$

$$0.399 \times 0.659 \times 10^{25} \frac{p}{\sqrt{M_r T}} \propto 2.63 \times 10^{24} \frac{p}{\sqrt{M_r T}}$$

$$R \propto \frac{2.63 \times 10^{24}}{7.5 \times 10^{23} \times 10^4} \frac{p}{\sqrt{M_r T}} \propto 3.5 \times 10^{22} \frac{p}{\sqrt{M_r T}}$$

k_B = Boltzmann's constant (J/K)

T = Temperature (K)

p = Pressure (torr)

$$R \propto 3.5 \times 10^{22} \frac{p}{\sqrt{M_r T}} \text{ molecule s}^{-1} \text{ cm}^2$$

$$M_{O_2} = 32$$

O_2 at $p = 760$ torr, 293 K

$$R = 2.75 \times 10^{23} \text{ molecules s}^{-1} \text{ cm}^2$$

O_2 at $p = 1 \times 10^{-6}$ torr, 293 K

$$R = 3.61 \times 10^{14} \text{ molecules s}^{-1} \text{ cm}^2$$

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J Spring



Gas Kinetics:

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