

Exam.	B. E. (Computer Engineering)		
Level	BE	Full Marks	80
Programme	BCT	Pass Marks	32
Year / Part	III / II	Time	3 hrs.

Subject: -Operating System (CT 656)

- ✓ Candidates are required to give their answers in their own words as far as practicable.
- ✓ Attempt All questions.
- ✓ The figures in the margin indicate Full Marks.
- ✓ Assume suitable data if necessary.

1. What is the role of system call in an operating system? Write down the major differences between following types of operating system. [2+8]
- a) Batch system
 - b) Interactive system
 - c) Real time system
 - d) Time sharing system

2. State 5- State process model. From the given following information: [2+8]

Process	Burst Time	priority
P ₁	10	3
P ₂	1	1
P ₃	2	3
P ₄	1	4
P ₅	2	2

Assume the processes arrived in the order P₁, P₂, P₃, P₄ and P₅ all time 0, priority 1 as highest and 4 as lowest.

- a) Draw the gantt chart
 - b) Calculate average waiting time and average turnaround time for the following scheduling algorithm.
 - (i) Round robin (quantum= 1)
 - (ii) priority preemptive
 - (iii) preemptive SJF
 - (iv) FCFS
3. What is critical section problem? Why must the executing the critical section be mutually exclusive? Describe how race condition occurs in producer- consumer problem and use semaphore to solve it. [2+2+6]
4. Define page fault and demand paging. Consider logical address spaces of eight pages of 1024 words, each mapped onto a physical memory of 32 frames then, [4+6]
- a) How many bits are in logical address and physical address?
 - b) How paging will be done?
5. What is the role of file system? How can files be implemented? List any three of them with advantages and disadvantages of each. [2+6]

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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 $\text{NaOH (aq)} + \text{C}_6\text{H}_4\text{COOH (aq)} \rightarrow \text{C}_6\text{H}_4\text{COONa (aq)} + \text{H}_2\text{O (l)}$
the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the