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How To Become A Professional Bachelor

Jin-Ying Zhang



How To Become A Professional Bachelor:

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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} + \text{MnO}_4^- + 4\text{H}^+$ is

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stoichiometric coefficient of H_2SO_4 if the above equation is balanced in an acidic medium stuck use a hint report a problem

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stages first bromine reacts with OH^- ions to form bromide ions and BrO^- ions second the BrO^- ions react to form bromide ions and bromate BrO_3^- ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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