

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, from  $K_w$  and  $K_a$  or  $K_b$  for conjugate acid-base pairs.
2. Calculate the pH of weak acid and weak base solutions.

## Vocabulary

- acid dissociation constant
- base dissociation constant

- conjugate acid-base pair
- hydrolysis
- indicator

- neutralization
- polyprotic acid

## Key Equations

$$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \quad \text{Equation 19.1}$$

$$K_b = \frac{[\text{B}][\text{H}_3\text{O}^+]}{[\text{BH}^+]} \quad \text{Equation 19.2}$$

## Part A Completion

Use the appropriate  $K_a$  or  $K_b$  value to identify the conjugate acid or base of the following substances. Write the chemical formula of the conjugate acid or base in the space provided.

1.  $\text{H}_2\text{SO}_4$  \_\_\_\_\_
2.  $\text{H}_2\text{PO}_4^-$  \_\_\_\_\_
3.  $\text{H}_2\text{O}$  \_\_\_\_\_
4.  $\text{H}_2\text{S}$  \_\_\_\_\_
5.  $\text{H}_2\text{SO}_3$  \_\_\_\_\_
6.  $\text{H}_2\text{SO}_4$  \_\_\_\_\_
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