

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) volume, molarity, pH, concentration of ions, amount of substance

(A) $V = \text{dm}^3$, $M = \text{mol dm}^{-3}$

(B) $p = 10^{-\text{pH}}$, $V = \text{dm}^3$, $M = \text{mol dm}^{-3}$, $n = \text{mol}$
 (C) $V = \text{dm}^3$, $p = 10^{-\text{pH}}$, $M = \text{mol dm}^{-3}$, $n = \text{mol}$ (not all the same)

2. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, 10% of the indicator is in its protonated form. What is the pH of the solution? (correct)

(A) $\text{pH} = 4.5 + \log 10$ (B) $\text{pH} = 4.5 - \log 10$

(C) $\text{pH} = 4.5$ (D) $\text{pH} = 4.5 + \log 0.1$ (E) $\text{pH} = 4.5 - \log 0.1$

3. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, 10% of the indicator is in its protonated form. What is the concentration of the indicator? (correct)

(A) 0.1 (B) 0.01 (C) 0.001 (D) 0.0001

4. Which indicator is most appropriate for the following reaction? (correct)



(A) 1 (B) 2 (C) 3 (D) 4

5. A solution of acetic acid and its conjugate base is prepared with a molar ratio of 1:1. The pK_a of acetic acid is 4.75. What is the pH of the solution? (correct)



(A) 4.75 (B) 4.75 (C) 5.25 (D) 5.25 (E) 5.75

6. An acid-base equilibrium is shown below. What is the pK_a of the acid? (correct)



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