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# Hands In Exile

**RJ Shavelson**



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 $\text{H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g}) \rightarrow \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow$

co<sub>2</sub> g 3 h<sub>2</sub> o g Δh 283 kj g 1 c<sub>2</sub>h<sub>6</sub> g

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web thus for reaction 9 2 in which 1 mole of  $h_3o^+$  and 1 mole of  $oh^-$  combine to form 2 moles of  $h_2o$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{products} - \Delta h_{reactants}$  is positive if heat is absorbed that is if  $\Delta h_{products}$

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