

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|------------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium | 5. hydroxide ion (OH^-) | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (14.1) | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) |
| 3. $\text{pH} = -\log[\text{H}^+]$ (14.3) | 4. $\text{pOH} = -\log[\text{OH}^-]$ (14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the acid-base reaction shown in the following equation.

1. _____ Which product is the _____ relative to the conjugate acid-base pair.
2. _____ Which product is the _____ of the conjugate base of the acid which produced an H^+ ion in the reaction.
3. _____ Which product is the _____ form of _____.
4. _____ Which product is the _____ of the _____ ion.
5. _____ Which product is the _____ of _____.
6. _____ Which product is the _____ form of _____.
7. _____ Which product is the _____ of _____.
8. _____ Which product is the _____ of _____.

- | | |
|-----|-------|
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19 2 Hydrogen Ions And Acidity Section Review Answers

JA Banks



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