

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide ion concentrations, molar ratios, and pH values for weak acid-conjugate base systems.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. weak acid | 6. hydroxide ion |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{pH} = -\log[\text{H}^+]$ (Eq. 17.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 17.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 17.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 17.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the system. Also indicate which species acts as the conjugate acid.

1. H_2O and OH^- (conjugate acid-base pair)
2. H_2O and H_3O^+ (conjugate acid-base pair)
3. H_2O and H^+ (conjugate acid-base pair)
4. H_2O and H^+ (conjugate acid-base pair)
5. H_2O and H^+ (conjugate acid-base pair)
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18. H_2O and H^+ (conjugate acid-base pair)
19. H_2O and H^+ (conjugate acid-base pair)
20. H_2O and H^+ (conjugate acid-base pair)

1. H_2O and OH^-

2. H_2O and H_3O^+

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6. H_2O and H^+

7. H_2O and H^+

8. H_2O and H^+

9. H_2O and H^+

10. H_2O and H^+

11. H_2O and H^+

12. H_2O and H^+

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Puja Mehta



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