

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH⁻] and [H⁺].
3. Calculate the pH of weak (or strong) acids.

Vocabulary

- | | | |
|-------------------|---|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium | 5. hydronium ion (H ₃ O ⁺) | 6. autoionization |
| | 7. weak acid | 8. pH |

Key Equations

- | | |
|--------------------------------------|--------------------------------------|
| 1. $[H^+] = 10^{-pH}$ (see page 513) | 2. $pH = -\log [H^+]$ (see page 513) |
| 3. $pH = -\log [H^+]$ (see page 513) | 4. $[H^+] = 10^{-pH}$ (see page 513) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the _____ of a weak hydroxide ion and a hydronium ion? _____ of the conjugate acid-base pair.
2. _____ of a strong acid is the _____ of a weak acid.
3. _____ of a weak acid is the _____ of a strong acid.
4. _____ of a weak acid is the _____ of a strong acid.
5. _____ of a weak acid is the _____ of a strong acid.
6. _____ of a weak acid is the _____ of a strong acid.
7. _____ of a weak acid is the _____ of a strong acid.
8. _____ of a weak acid is the _____ of a strong acid.
9. _____ of a weak acid is the _____ of a strong acid.
10. _____ of a weak acid is the _____ of a strong acid.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

19 2 Hydrogen Ions And Acidity Answer Key

Y Pai

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red-to-white gradient semi-circle on the right side of the bar.

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