

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) 10^{-14} (B) 10^{-7}

2. In a 0.1 M solution of H_2SO_4 , $[\text{H}^+] = 0.11 \text{ M}$, $[\text{HSO}_4^-] = 0.09 \text{ M}$, $[\text{SO}_4^{2-}] = 0.01 \text{ M}$, $[\text{H}_2\text{O}] = 55.5 \text{ M}$, $[\text{H}_2\text{SO}_4] = 0.09 \text{ M}$.
 (A) $[\text{H}^+] = 0.11 \text{ M}$, $[\text{HSO}_4^-] = 0.09 \text{ M}$, $[\text{SO}_4^{2-}] = 0.01 \text{ M}$, $[\text{H}_2\text{O}] = 55.5 \text{ M}$, $[\text{H}_2\text{SO}_4] = 0.09 \text{ M}$
 (B) $[\text{H}^+] = 0.11 \text{ M}$, $[\text{HSO}_4^-] = 0.09 \text{ M}$, $[\text{SO}_4^{2-}] = 0.01 \text{ M}$, $[\text{H}_2\text{O}] = 55.5 \text{ M}$, $[\text{H}_2\text{SO}_4] = 0.09 \text{ M}$

3. In a 0.1 M solution of H_2SO_4 , $[\text{H}^+] = 0.11 \text{ M}$, $[\text{HSO}_4^-] = 0.09 \text{ M}$, $[\text{SO}_4^{2-}] = 0.01 \text{ M}$, $[\text{H}_2\text{O}] = 55.5 \text{ M}$, $[\text{H}_2\text{SO}_4] = 0.09 \text{ M}$.
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4. In a 0.1 M solution of H_2SO_4 , $[\text{H}^+] = 0.11 \text{ M}$, $[\text{HSO}_4^-] = 0.09 \text{ M}$, $[\text{SO}_4^{2-}] = 0.01 \text{ M}$, $[\text{H}_2\text{O}] = 55.5 \text{ M}$, $[\text{H}_2\text{SO}_4] = 0.09 \text{ M}$.

5. Which is the correct chemical equation for the reaction of H_2SO_4 with NaOH ? (correct)
- (A) $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O}$
 (B) $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
 (C) $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
 (D) $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O}$

6. Which is the correct chemical equation for the reaction of H_2SO_4 with NaOH ? (correct)
- (A) $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O}$
 (B) $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
 (C) $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
 (D) $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O}$

7. Which is the correct chemical equation for the following reaction? (correct)



- (A) $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O}$
 (B) $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
 (C) $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
 (D) $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O}$

8. A solution of H_2SO_4 and NaOH are mixed together and the pH is measured. If the pH is 12.0, which of the following is the correct chemical equation for the reaction? (correct)



- (C) $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$ (D) $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O}$

9. In a 0.1 M solution of H_2SO_4 , $[\text{H}^+] = 0.11 \text{ M}$, $[\text{HSO}_4^-] = 0.09 \text{ M}$, $[\text{SO}_4^{2-}] = 0.01 \text{ M}$, $[\text{H}_2\text{O}] = 55.5 \text{ M}$, $[\text{H}_2\text{SO}_4] = 0.09 \text{ M}$.
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(A) 10^{-14} (B) 10^{-7}

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Paul Foglino, Princeton Review (Firm)

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Table of Contents Multiple Choice Stoichiometry Test

1. Understanding the eBook Multiple Choice Stoichiometry Test
 - The Rise of Digital Reading Multiple Choice Stoichiometry Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Stoichiometry Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Stoichiometry Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Stoichiometry Test
 - Personalized Recommendations
 - Multiple Choice Stoichiometry Test User Reviews and Ratings
 - Multiple Choice Stoichiometry Test and Bestseller Lists
5. Accessing Multiple Choice Stoichiometry Test Free and Paid eBooks
 - Multiple Choice Stoichiometry Test Public Domain eBooks
 - Multiple Choice Stoichiometry Test eBook Subscription Services

- Multiple Choice Stoichiometry Test Budget-Friendly Options
- 6. Navigating Multiple Choice Stoichiometry Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Multiple Choice Stoichiometry Test Compatibility with Devices
 - Multiple Choice Stoichiometry Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Stoichiometry Test
 - Highlighting and Note-Taking Multiple Choice Stoichiometry Test
 - Interactive Elements Multiple Choice Stoichiometry Test
- 8. Staying Engaged with Multiple Choice Stoichiometry Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Stoichiometry Test
- 9. Balancing eBooks and Physical Books Multiple Choice Stoichiometry Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Stoichiometry Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Choice Stoichiometry Test
 - Setting Reading Goals Multiple Choice Stoichiometry Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Choice Stoichiometry Test
 - Fact-Checking eBook Content of Multiple Choice Stoichiometry Test
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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