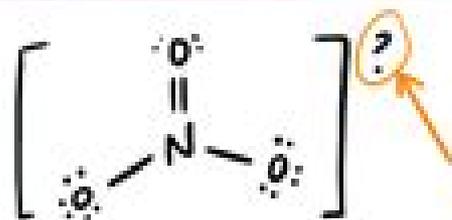


[FORMAL CHARGE]^{+/-}

$$F.C. = (\text{valence } e^-) - (\text{nonbonding } e^-) - \frac{1}{2} (\text{bonding } e^-)$$

- * The amount of electrons in a specific atom when the molecule shares electrons equally.
- * Only applies to covalent molecules.

Examples



Overall charge comes from addition of all atomic formal charges!

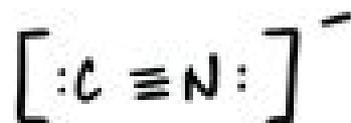
Q: What is the formal charge of N?

A:

$$F.C. = 5 - 0 - \left(\frac{1}{2} \times 8\right)$$

$$F.C. = 5 - 4$$

$$F.C. = +1$$



Q: What is the formal charge of carbon?

$$A: F.C. = 4 - 2 - \left(\frac{1}{2} \times 6\right)$$

$$F.C. = 4 - 2 - 3$$

$$F.C. = -1$$

Formal Charges

Camilla Rothe



Formal Charges:

The Enigmatic Realm of **Formal Charges**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing in short supply of extraordinary. Within the captivating pages of **Formal Charges** a literary masterpiece penned with a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting affect on the hearts and minds of those who partake in its reading experience.

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