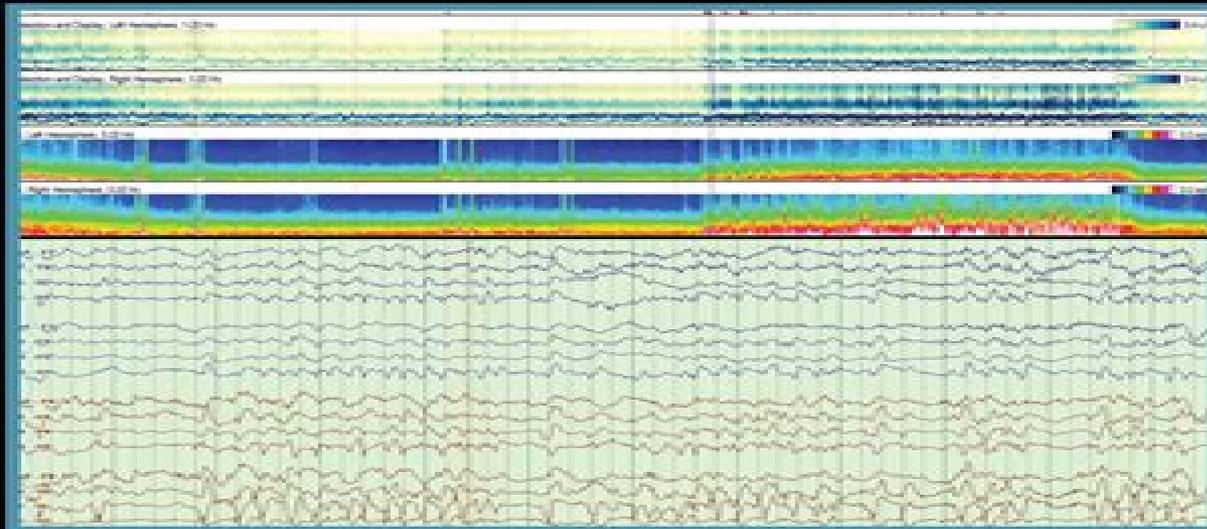


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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \text{ aq } 2\text{NaOH aq Na}_2\text{C}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction $\text{NaOH aq C}_6\text{H}_4\text{COOH aq C}_6\text{H}_4\text{COONa aq H}_2\text{O l}$ the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the

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dextrose and or lipids tpn refers to amino acids dextrose and fat emulsion if a drug is stated compatible in tpn it

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