

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) volume of precipitate, concentration of ions, molarity and amount

(A) $V = \frac{m}{\rho}$ (B) $C = \frac{m}{V}$

(C) $pH = -\log_{10} \frac{m}{V}$ (D) $pH = -\log_{10} \frac{m}{V}$ (E) $pH = -\log_{10} \frac{m}{V}$ (F) $pH = -\log_{10} \frac{m}{V}$

2. An ionic solution contains sodium hydroxide (solute) and water (solvent). If 1.0 g of NaOH is added to 100 g of water, what is the pH of the solution? (correct)

(A) $pH = 14$ (B) $pH = 13$ (C) $pH = 12$ (D) $pH = 11$ (E) $pH = 10$

(F) $pH = 9$ (G) $pH = 8$ (H) $pH = 7$ (I) $pH = 6$ (J) $pH = 5$

3. An ionic solution contains NaOH (solute) and water (solvent). What is the pH of the solution if 1.0 g of NaOH is added to 100 g of water? (correct)

(A) 14 (B) 13 (C) 12 (D) 11

4. Which is the correct combination of properties for this following equation? (correct)



(A) exothermic, redox (B) exothermic, not redox
 (C) endothermic, redox (D) endothermic, not redox

5. A solution contains 1.0 g of NaOH and 100 g of water. The pH of the solution is 13.0. If 1.0 g of NaOH is added to the solution, what is the pH of the solution? (correct)

(A) 14.0 (B) 13.0 (C) 12.0 (D) 11.0 (E) 10.0

(F) 9.0 (G) 8.0 (H) 7.0 (I) 6.0 (J) 5.0

6. An acid-base reaction is shown below. What is the pH of the solution after the reaction is complete? (correct)



Multiple Choice Stoichiometry Test

Brendan G. Carr



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mélangez avec un fouet couvrez et laissez reposer pendant environ 40 minutes

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