

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

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Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1V_1=P_2V_2$
Isolate for P₂: $P_2=(P_1V_1)/V_2$
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1=34.75 + 273=307.75 \text{ K}$
 $T_2=78.41 + 273=351.41 \text{ K}$
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
Get rid of P
 $V_1/T_1=V_2/T_2$
Isolate V₂: $V_2=(V_1/T_1)T_2$
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV=nRT$ will work. (n= moles and R= 0.08206 L atm/K mol)
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$ OOPS we need to convert from grams to moles!, $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$
n: Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$
Isolate V: $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

2nd Semester Final Review Guide Chemistry

Puja Mehta



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Table of Contents 2nd Semester Final Review Guide Chemistry

1. Understanding the eBook 2nd Semester Final Review Guide Chemistry
 - The Rise of Digital Reading 2nd Semester Final Review Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Final Review Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2nd Semester Final Review Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Final Review Guide Chemistry
 - Personalized Recommendations
 - 2nd Semester Final Review Guide Chemistry User Reviews and Ratings
 - 2nd Semester Final Review Guide Chemistry and Bestseller Lists
5. Accessing 2nd Semester Final Review Guide Chemistry Free and Paid eBooks
 - 2nd Semester Final Review Guide Chemistry Public Domain eBooks
 - 2nd Semester Final Review Guide Chemistry eBook Subscription Services
 - 2nd Semester Final Review Guide Chemistry Budget-Friendly Options

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 - 2nd Semester Final Review Guide Chemistry Compatibility with Devices
 - 2nd Semester Final Review Guide Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2nd Semester Final Review Guide Chemistry
 - Highlighting and Note-Taking 2nd Semester Final Review Guide Chemistry
 - Interactive Elements 2nd Semester Final Review Guide Chemistry
8. Staying Engaged with 2nd Semester Final Review Guide Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2nd Semester Final Review Guide Chemistry
9. Balancing eBooks and Physical Books 2nd Semester Final Review Guide Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2nd Semester Final Review Guide Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2nd Semester Final Review Guide Chemistry
 - Setting Reading Goals 2nd Semester Final Review Guide Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2nd Semester Final Review Guide Chemistry
 - Fact-Checking eBook Content of 2nd Semester Final Review Guide Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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