

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) volume, molarity, pH, concentration of ions, concentration of substance

(A)  $V = \text{dm}^3$ ,  $M = \text{mol dm}^{-3}$

(B)  $p = 10^{-\text{pH}}$ ,  $V = \text{dm}^3$ ,  $M = \text{mol dm}^{-3}$ ,  $n = \text{mol}$   
 (C)  $V = \text{dm}^3$ ,  $p = 10^{-\text{pH}}$ ,  $M = \text{mol dm}^{-3}$ ,  $n = \text{mol}$  (not all the same)

2. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of the indicator, 10% of the indicator is in its protonated form. What is the pH of the solution? (correct)

(A)  $\text{pH} = 4.5 + \log 10$  (B)  $\text{pH} = 4.5 - \log 10$

(C)  $\text{pH} = 4.5$  (D)  $\text{pH} = 4.5 + \log 0.1$  (E)  $\text{pH} = 4.5 - \log 0.1$

3. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of the indicator, 10% of the indicator is in its protonated form. What is the concentration of the indicator? (correct)

(A) 0.1 (B) 1.0 (C) 0.01 (D) 0.10 (E) 0.001

4. Which indicator is most appropriate for the following reaction? (correct)



(A) 1 (B) 2 (C) 3 (D) 4 (E) 5

5. Amino acids react with HCl to form zwitterions. The zwitterion form of the amino acid glycine is shown. If a 0.10 M solution of glycine is prepared and the pH is 9.0, what mass of HCl must be added? (correct)



(A) 0.010 g (B) 0.020 g (C) 0.040 g (D) 0.080 g (E) 0.10 g

6. An acid-base equilibrium,  $\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{OH}^-$ , is shown. What is the  $\text{pK}_a$  of the acid  $\text{H}_2\text{CO}_3$ ? (correct)



# Multiple Choice Stoichiometry Test

**Emilie Sanchez**

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