

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) 10^{-14} (B) 10^{-7} (C) 10^{-1} (D) 10^{-1}

(A) 10^{-14} (B) 10^{-7} (C) 10^{-1} (D) 10^{-1}
 (E) 10^{-14} (F) 10^{-7} (G) 10^{-1} (H) 10^{-1} (I) 10^{-14} (J) 10^{-7} (K) 10^{-1} (L) 10^{-1}

2. An acid-base reaction between two weak acids and their conjugate bases, $HA_1 + A_2 \rightleftharpoons A_1 + HA_2$, is shown below. The equilibrium constant for this reaction is K . What is the value of K ? (correct)

(A) 10^{-14} (B) 10^{-7} (C) 10^{-1} (D) 10^{-1} (E) 10^{-14} (F) 10^{-7} (G) 10^{-1} (H) 10^{-1} (I) 10^{-14} (J) 10^{-7} (K) 10^{-1} (L) 10^{-1}

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3. An acid-base reaction between two weak acids and their conjugate bases, $HA_1 + A_2 \rightleftharpoons A_1 + HA_2$, is shown below. The equilibrium constant for this reaction is K . What is the value of K ? (correct)

(A) 10^{-14} (B) 10^{-7} (C) 10^{-1} (D) 10^{-1} (E) 10^{-14} (F) 10^{-7} (G) 10^{-1} (H) 10^{-1} (I) 10^{-14} (J) 10^{-7} (K) 10^{-1} (L) 10^{-1}

4. Which of the following is a weak acid? (correct)



(A) 10^{-14} (B) 10^{-7} (C) 10^{-1} (D) 10^{-1} (E) 10^{-14} (F) 10^{-7} (G) 10^{-1} (H) 10^{-1} (I) 10^{-14} (J) 10^{-7} (K) 10^{-1} (L) 10^{-1}

5. A weak acid, HA , is shown below. The equilibrium constant for the reaction $HA + H_2O \rightleftharpoons A^- + H_3O^+$ is K . What is the value of K ? (correct)



(A) 10^{-14} (B) 10^{-7} (C) 10^{-1} (D) 10^{-1} (E) 10^{-14} (F) 10^{-7} (G) 10^{-1} (H) 10^{-1} (I) 10^{-14} (J) 10^{-7} (K) 10^{-1} (L) 10^{-1}

6. An acid-base reaction between two weak acids and their conjugate bases, $HA_1 + A_2 \rightleftharpoons A_1 + HA_2$, is shown below. The equilibrium constant for this reaction is K . What is the value of K ? (correct)



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