

18.4 Molecular Structure: Mass

Multiple Choice

1. The molecular formula of a compound is the chemical formula written _____
 (A) percent composition (B) percent yield
 (C) molar formula (D) structural formula
2. Percent of hydrogen in $\text{C}_2\text{H}_6\text{O}$ is _____
 (A) 15.0% (B) 13.3%

- (C) 16.0% (D) 14.3% (E) 17.0% (F) none of these

3. An organic compound contains oxygen, carbon, and hydrogen. A 0.100 g sample contained 0.0273 g hydrogen, and only produced 0.0441 g CO_2 when the sample was burned.



- (F) $\text{C}_2\text{H}_4\text{O}_3$ (G) $\text{C}_2\text{H}_2\text{O}_3$ (H) $\text{C}_2\text{H}_4\text{O}_4$ (I) $\text{C}_2\text{H}_2\text{O}_4$ (J) none of these

4. An organic compound contains oxygen, carbon, and hydrogen. A 0.100 g sample contained 0.0273 g hydrogen, and only produced 0.0441 g CO_2 when the sample was burned.

- (A) C_2H_6 (B) C_2H_4 (C) C_2H_2 (D) $\text{C}_2\text{H}_4\text{O}$ (E) $\text{C}_2\text{H}_2\text{O}$

5. Percent carbon in C_2H_6 is _____
 (A) 80.0% (B) 85.7% (C) 83.3% (D) 82.0%



- (I) $\text{C}_2\text{H}_4\text{O}_2$ (J) $\text{C}_2\text{H}_2\text{O}_2$ (K) $\text{C}_2\text{H}_4\text{O}_3$ (L) $\text{C}_2\text{H}_2\text{O}_3$

6. A 0.100 g sample of an organic compound with molar mass 100.0 g/mol contains 0.0400 g hydrogen, 0.0800 g carbon, and 0.0800 g oxygen. The sample was burned, and 0.0800 g CO_2 was produced.



- (F) $\text{C}_2\text{H}_2\text{O}_3$ (G) $\text{C}_2\text{H}_4\text{O}_4$ (H) $\text{C}_2\text{H}_2\text{O}_4$ (I) $\text{C}_2\text{H}_4\text{O}_5$ (J) $\text{C}_2\text{H}_2\text{O}_5$

7. An organic compound, $\text{C}_2\text{H}_4\text{O}_2$, was burned, and 0.0800 g CO_2 was produced. The sample was 0.100 g. The sample was burned, and 0.0800 g CO_2 was produced. The sample was 0.100 g.



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David Baud



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