

Chapters 1-3 Notes

11 Pages

DIMENSIONAL ANALYSIS

1. Write your starting unit (what you know) from the problem?
2. Write your desired unit (what you want) to change the starting unit?
• usually want mass unit
3. You should be left with the desired unit after calculations
• looking units may be used depending on the problem

Starting unit \times Conversion = desired unit
Starting unit

Ex: (each component) \times (conversion factor) = (unit)
Ex: $100 \text{ g} \times \frac{1 \text{ kg}}{1000 \text{ g}} = 0.1 \text{ kg}$

molecular mass: sum of the atomic masses (amu) of molecules

$\text{molecular mass (amu)} = \sum (\text{number of atoms}) (\text{atomic mass})$

Formula mass: sum of atomic masses (amu) in a formula unit of ionic compound

$\text{Formula mass (amu)} = \sum (\text{number of atoms}) (\text{atomic mass})$

Solution

Mass % = $\frac{\text{mass of solute}}{\text{mass of solution}} \times 100\%$

Ex: 25.0% Cu = 25.0 g Cu / 100 g soln
25.0% Cu = 25.0 g Cu / 100 g soln

VS Empirical Formula

Empirical formula: simplest whole number ratio of atoms in a molecule

Ex: H_2O and $\text{C}_2\text{H}_4\text{O}$

Chemical nomenclature

Binary compounds: two elements

Ex: FeCl_2 (Iron(II) chloride)

Ex: Cu_2S (Copper(I) sulfide)

general chemistry

Acid

Substance that **gives** H^+ when dissolved in water

Ex: $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$

BASE

Substance that **accepts** H^+

Ex: NaOH , Ba(OH)_2

hydro

prefix for acids

Ex: H_3PO_4 (phosphoric acid)

naming

1. When an H is present, name the anion with "acid"
2. When an H is not present, name the anion with "ate"
3. Indicate # of H's present when not all H's are present

Ex: H_3PO_4 (phosphoric acid)
Ex: HPO_4^{2-} (hydrogen phosphate)

BALANCING CHEMICAL REACTIONS

Change coefficients (numbers in front of molecules) to equal number of atoms of each element on both sides of equation

Ex: $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

Element	Reactants	Products
C	2	2
H	6	6
O	2	4

Try to make numbers equal

Ex: $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$

PERCENT YIELD

Calculated from actual yield and theoretical yield

$\% \text{ yield} = \frac{\text{Actual yield}}{\text{Theoretical yield}} \times 100\%$

Ex: $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$

Ex: $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$

General Chemistry 1 Lecture Notes And Workbook

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