

14.5

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH and pOH from the concentration of either ion.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. hydronium ion | 6. hydronium ion |
| | 7. weak acid | 8. pH |

Key Equations

- | | |
|---|---|
| 1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 14.5) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.6) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.7) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.8) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in (a). Label the acid and base members with their conjugate partners.

1. H_2O and OH^- (acid and base)
2. H_2O and H_3O^+ (acid and base)
3. H_2O and H^+ (acid and base)
4. H_2O and H_2O (acid and base)
5. H_2O and H_2O (acid and base)
6. H_2O and H_2O (acid and base)
7. H_2O and H_2O (acid and base)
8. H_2O and H_2O (acid and base)
9. H_2O and H_2O (acid and base)
10. H_2O and H_2O (acid and base)
11. H_2O and H_2O (acid and base)
12. H_2O and H_2O (acid and base)
13. H_2O and H_2O (acid and base)
14. H_2O and H_2O (acid and base)
15. H_2O and H_2O (acid and base)
16. H_2O and H_2O (acid and base)
17. H_2O and H_2O (acid and base)
18. H_2O and H_2O (acid and base)
19. H_2O and H_2O (acid and base)
20. H_2O and H_2O (acid and base)

21. H_2O and H_2O

22. H_2O and H_2O

23. H_2O and H_2O

24. H_2O and H_2O

25. H_2O and H_2O

26. H_2O and H_2O

27. H_2O and H_2O

28. H_2O and H_2O

29. H_2O and H_2O

30. H_2O and H_2O

31. H_2O and H_2O

32. H_2O and H_2O

19 2 Hydrogen Ions And Acidity Worksheet Answers

F Rizvi



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