

18.4 Acid-Base Chemistry: Theory

Multiple-Choice:

1. The acid-dissociation constant (K_a) of a weak acid is 1.0×10^{-4} . The pK_a of this acid is _____.

(A) 4.00 (B) 10.00 (C) 1.00 (D) 0.00

2. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

3. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

4. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

5. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

6. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

7. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

8. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

9. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

10. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

11. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

12. The pH of a 0.10 M solution of a weak acid is 3.00. The K_a of this acid is _____.

(A) 1.0×10^{-3} (B) 1.0×10^{-4} (C) 1.0×10^{-5} (D) 1.0×10^{-6}

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KJ Lindholm-Leary



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