

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $10^{-14}$  (B)  $10^{-7}$

2. In a 0.1 mol/L acetic acid solution, which of the following is correct? (correct)
- (A)  $[H^+] = 10^{-7}$  mol/L (B)  $[H^+] = 10^{-8}$  mol/L (C)  $[H^+] = 10^{-9}$  mol/L (D)  $[H^+] = 10^{-10}$  mol/L

3. In a 0.1 mol/L acetic acid solution, which of the following is correct? (correct)
- (A)  $[H^+] = 10^{-7}$  mol/L (B)  $[H^+] = 10^{-8}$  mol/L (C)  $[H^+] = 10^{-9}$  mol/L (D)  $[H^+] = 10^{-10}$  mol/L

(A)  $10^{-7}$  (B)  $10^{-8}$  (C)  $10^{-9}$  (D)  $10^{-10}$

4. In a 0.1 mol/L acetic acid solution, which of the following is correct? (correct)
- (A)  $[H^+] = 10^{-7}$  mol/L (B)  $[H^+] = 10^{-8}$  mol/L (C)  $[H^+] = 10^{-9}$  mol/L (D)  $[H^+] = 10^{-10}$  mol/L

5. In a 0.1 mol/L acetic acid solution, which of the following is correct? (correct)
- (A)  $[H^+] = 10^{-7}$  mol/L (B)  $[H^+] = 10^{-8}$  mol/L (C)  $[H^+] = 10^{-9}$  mol/L (D)  $[H^+] = 10^{-10}$  mol/L

(A)  $10^{-7}$  (B)  $10^{-8}$  (C)  $10^{-9}$  (D)  $10^{-10}$

6. Which of the following is correct? (correct)
- (A)  $[H^+] = 10^{-7}$  mol/L (B)  $[H^+] = 10^{-8}$  mol/L (C)  $[H^+] = 10^{-9}$  mol/L (D)  $[H^+] = 10^{-10}$  mol/L

(A)  $10^{-7}$  (B)  $10^{-8}$  (C)  $10^{-9}$  (D)  $10^{-10}$

7. In a 0.1 mol/L acetic acid solution, which of the following is correct? (correct)
- (A)  $[H^+] = 10^{-7}$  mol/L (B)  $[H^+] = 10^{-8}$  mol/L (C)  $[H^+] = 10^{-9}$  mol/L (D)  $[H^+] = 10^{-10}$  mol/L

(A)  $10^{-7}$  (B)  $10^{-8}$  (C)  $10^{-9}$  (D)  $10^{-10}$

8. In a 0.1 mol/L acetic acid solution, which of the following is correct? (correct)
- (A)  $[H^+] = 10^{-7}$  mol/L (B)  $[H^+] = 10^{-8}$  mol/L (C)  $[H^+] = 10^{-9}$  mol/L (D)  $[H^+] = 10^{-10}$  mol/L

9. In a 0.1 mol/L acetic acid solution, which of the following is correct? (correct)
- (A)  $[H^+] = 10^{-7}$  mol/L (B)  $[H^+] = 10^{-8}$  mol/L (C)  $[H^+] = 10^{-9}$  mol/L (D)  $[H^+] = 10^{-10}$  mol/L

(A)  $10^{-7}$  (B)  $10^{-8}$  (C)  $10^{-9}$  (D)  $10^{-10}$

# Multiple Choice Stoichiometry Test

**L Cohen**



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