

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|---------------------------------------|--------------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. acid ionization constant (K_a) | 6. hydronium ion concentration |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|------------------------------------------------------------------------------|----------------------------------------------------------------------------|
| 1. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ (Eq. 17.1) | 2. $K_b = \frac{[\text{B}][\text{H}_2\text{O}]}{[\text{BH}^+]}$ (Eq. 17.2) |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 17.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 17.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the reaction. Write the number in the space provided.

1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (acid-base autoionization)
2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (acid-base autoionization)
3. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (acid-base autoionization)
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8. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (acid-base autoionization)
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10. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (acid-base autoionization)

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Emilie Sanchez



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