

18.4 Molecular Structure: Mass

Multiple Choice

1. The molecular formula of a compound is the chemical formula written _____
 (A) percent composition (B) percent yield
 (C) molar mass (D) structural analysis
2. Percent of hydrogen in $\text{C}_2\text{H}_6\text{O}$ is _____
 (A) 15.0% (B) 12.5%

3. In a gas the other reactions represent oxygen mixed with nitrogen, 2.5 g of gas was produced by the reaction, and only produced 1.5 g of gas. What is the percent yield of sulfur in this reaction?



(A) 60% (B) 75% (C) 80% (D) 90% (E) 100%

4. In this reaction $\text{ZnCO}_3 \rightarrow \text{ZnO} + \text{CO}_2$, what is the percent yield of oxygen?
 (A) 10% (B) 20% (C) 30% (D) 40%

5. What is the percent yield of oxygen in the following reaction?



(A) 10% (B) 20% (C) 30% (D) 40%

6. Aluminum reacts with H_2 to produce hydrogen gas. If 1.5 g of aluminum is used and 0.5 g of hydrogen gas is produced, what is the percent yield?



(A) 10% (B) 20% (C) 30% (D) 40% (E) 50%

7. In a full experiment, structural formula of $\text{C}_2\text{H}_4\text{O}_2$ was given and solution CH_3COOH after the reaction was completed, making a gas product with 1.5 g of gas in the reaction mixture. What is the percent yield?



Multiple Choice Stoichiometry Test

D Keegan

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular glow behind it.

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