

Chemistry: 2nd Semester Exam

Select the correct oxidation numbers for the following elements.

- Oxygen in O_2
a. -2 b. +2 c. -1 d. +1 e. 0
- Sodium in $NaCl$
a. -2 b. +2 c. -1 d. +1 e. 0
- Phosphorus in H_3PO_4
a. -5 b. -3 c. 0 d. +3 e. +5
- H in H_2O
a. -2 b. +2 c. -1 d. +1 e. 0
- O in H_2O
a. -2 b. +2 c. -1 d. +1 e. 0
- In which substance does bromine have an oxidation number of +1?
a. Br_2 b. HBr c. $HBrO$ d. $HBrO_2$
- Which species is the oxidizing agent in the following reaction? $Cl_2(aq) + 2I^-(aq) \rightarrow I_2(aq) + 2Cl^-(aq)$
a. Cl_2 b. $2I^-$ c. I_2 d. $2Cl^-$
- Consider the following: $2MnO_4^- + 5H_2SO_3 \rightarrow 2Mn^{2+} + 3H_2O + 5SO_4^{2-} + 4H^+$
Which of the following elements does not experience a change in oxidation number?
a. S in H_2SO_3
b. O in MnO_4^-
c. Mn in MnO_4^-
- In which of the following unbalanced equations does chromium undergo oxidation?
a. $Cr^{3+} \rightarrow Cr$
b. $Cr^{2+} \rightarrow Cr^{3+}$
c. $Cr^{3+} \rightarrow Cr_2O_7^{2-}$
d. $CrO_4^{2-} \rightarrow Cr_2O_7^{2-}$
- In a redox reaction, the species which loses electrons is _____.
a. reduced b. oxidized c. neutralized
- Consider the following: $Cl_2 \rightarrow ClO_2^-$
The balanced half reaction is:
a. $Cl_2 + 3H_2O \rightarrow ClO_2^- + 6H^+ + 5e^-$
b. $Cl_2 + 3H_2O \rightarrow 2ClO_2^- + 6H^+ + 4e^-$
c. $Cl_2 + 6H_2O + 2e^- \rightarrow 2ClO_2^- + 6H_2$
d. $Cl_2 + 6H_2O \rightarrow 2ClO_2^- + 12H^+ + 10e^-$

Label the Acid, Base, Conjugate Acid, and Conjugate Base



- $HC_2H_3O_2$ serves as the
a. Acid b. Base c. Conjugate Acid d. Conjugate Base
- H_2O serves as the
a. Acid b. Base c. Conjugate Acid d. Conjugate Base
- H_3O^+ serves as the
a. Acid b. Base c. Conjugate Acid d. Conjugate Base
- $C_2H_3O_2^-$ serves as the
a. Acid b. Base c. Conjugate Acid d. Conjugate Base

2nd Semester Exam Chemistry Form

Scott C. Dulebohn



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