

16.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide ion concentration, hydronium ion, and pH, and use measured hydronium ion concentration to calculate pH.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. hydronium ion |
| 4. autoionization | 5. hydronium ion | 6. hydronium ion |
| | 7. hydronium ion | 7. pH |

Key Equations

- | | |
|--|--|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ | 4. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of 1.0000000000000000.

1. What product is the _____ of a weak hydroxide ion and a hydronium ion? _____ of the conjugate acid-base pair.
2. _____ of a weak acid and a hydroxide ion.
3. _____ of a weak acid and a hydroxide ion.
4. _____ of a weak acid and a hydroxide ion.
5. _____ of a weak acid and a hydroxide ion.
6. _____ of a weak acid and a hydroxide ion.
7. _____ of a weak acid and a hydroxide ion.
8. _____ of a weak acid and a hydroxide ion.
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15. _____ of a weak acid and a hydroxide ion.
16. _____ of a weak acid and a hydroxide ion.
17. _____ of a weak acid and a hydroxide ion.
18. _____ of a weak acid and a hydroxide ion.
19. _____ of a weak acid and a hydroxide ion.
20. _____ of a weak acid and a hydroxide ion.

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20. _____

19 2 Hydrogen Ions And Acidity Answers

**Thirumalaisamy P.
Velavancorresponding**



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