

## Answer Sheet for Sample Problems for Chemistry Level 1 Final Exam 2013-2014 Study Guide

### Electrons in Atoms – Chapter 5

1. What is the frequency of green light, which has a wavelength of  $4.90 \times 10^{-7} \text{ m}$ ?

$$c = \lambda \nu \quad \lambda = 4.90 \times 10^{-7} \text{ m}, c = 3.00 \times 10^8 \text{ m/s} \quad \nu = ? \text{ Hz} \quad \nu = \frac{c}{\lambda} = \frac{3.00 \times 10^8 \text{ m/s}}{4.90 \times 10^{-7} \text{ m}}$$

$$\underline{6.12 \times 10^{14} \text{ Hz}}$$

2. An x-ray has a wavelength of  $1.15 \times 10^{-10} \text{ m}$ . What is its frequency?

$$c = \lambda \nu \quad \lambda = 1.15 \times 10^{-10} \text{ m}, c = 3.00 \times 10^8 \text{ m/s} \quad \nu = ? \text{ Hz} \quad \nu = \frac{c}{\lambda} = \frac{3.00 \times 10^8 \text{ m/s}}{1.15 \times 10^{-10} \text{ m}}$$

$$\underline{2.61 \times 10^{18} \text{ Hz}}$$

3. Which type of electromagnetic radiation has the most energy: Microwaves, X-Rays, or Visible light? Why?

**Ans: X-Rays, b/c they have the shortest wavelength and highest frequency of the types listed and frequency and energy are directly related.**

4. What is the energy of a  $9.50 \times 10^{13} \text{ Hz}$  radiation?

$$E = h\nu \quad \nu = 9.50 \times 10^{13} \text{ Hz}, h = 6.626 \times 10^{-34} \text{ J s} \quad E = ? \text{ J}$$

$$E = h\nu = (6.626 \times 10^{-34} \text{ J s}) (9.50 \times 10^{13} \text{ s}^{-1}) = \underline{6.29 \times 10^{-19} \text{ J}}$$

5. The element with the electron configuration of  $[\text{Kr}]5s^24d^{10}5p^1$  is indium (In)

6. Barium (Ba)

7. Write out the electron configurations for the following elements: (For the exam make sure you know how to do both the complete and the noble gas short cut for electron configurations)

a. Tc                      **Ans:  $1s^22s^22p^63s^23p^44s^23d^{10}4p^65s^24d^5$  or  $[\text{Kr}]5s^24d^5$**

b. Aluminum            **Ans:  $1s^22s^22p^63s^23p^1$  or  $[\text{Ne}] 3s^23p^1$**

8. Why are Cr and Cu exceptions to writing the electron configurations of elements?

**They are exceptions because instead of following the pattern of filling up the d orbitals Cr actually has a half filled s and d orbitals and Cu has a half filled s and completely filled d orbitals. The reason why this occurs is because the electrons in the d-sublevel are more stable with this electron configuration than the expected configuration.**

9. a. Ca

b. As

c. Sn

a. 2    Ca.

b. 5    .As.

c. 4    . Sn .

### Periodic Table and Periodic Trends – Chapter 6

1. Compare & contrast how Mendeleev's periodic table was organized to how Moseley organized it.

# 2nd Semester Exam Review Sheet Answer

**David Baud**



**2nd Semester Exam Review Sheet Answer:**

## **2nd Semester Exam Review Sheet Answer** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**2nd Semester Exam Review Sheet Answer**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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