

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) 10^{-14} (B) 10^{-7}

2. In a 0.1 M solution of H_2SO_4 , $[\text{H}^+] = 0.11 \text{ M}$, $[\text{HSO}_4^-] = 0.09 \text{ M}$, $[\text{SO}_4^{2-}] = 0.01 \text{ M}$, $[\text{H}_2\text{O}] = 55.5 \text{ M}$, $[\text{OH}^-] = 10^{-14} \text{ M}$. What is the K_a of HSO_4^- ?
- (A) 1.1×10^{-2} (B) 1.1×10^{-3} (C) 1.1×10^{-4} (D) 1.1×10^{-5}

3. In a 0.1 M solution of acetic acid, CH_3COOH , $[\text{CH}_3\text{COOH}] = 0.09 \text{ M}$, $[\text{CH}_3\text{COO}^-] = 0.01 \text{ M}$, $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$. What is the K_a of CH_3COOH ?
- (A) 1.1×10^{-4} (B) 1.1×10^{-5} (C) 1.1×10^{-6} (D) 1.1×10^{-7}

(A) 1.1×10^{-4} (B) 1.1×10^{-5} (C) 1.1×10^{-6} (D) 1.1×10^{-7}

4. In a 0.1 M solution of H_2CO_3 , $[\text{H}_2\text{CO}_3] = 0.09 \text{ M}$, $[\text{HCO}_3^-] = 0.01 \text{ M}$, $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$. What is the K_a of H_2CO_3 ?
- (A) 1.1×10^{-4} (B) 1.1×10^{-5} (C) 1.1×10^{-6} (D) 1.1×10^{-7}

5. Which is the correct relationship between K_a and K_b for a weak acid and its conjugate base? (correct)
- (A) $K_a \times K_b = 1$ (B) $K_a \times K_b = 10^{-14}$ (C) $K_a \times K_b = 10^{-7}$ (D) $K_a \times K_b = 10^{-28}$

6. Which is the correct relationship between K_a and K_b for a weak acid and its conjugate base? (correct)
- (A) $K_a \times K_b = 1$ (B) $K_a \times K_b = 10^{-14}$ (C) $K_a \times K_b = 10^{-7}$ (D) $K_a \times K_b = 10^{-28}$

(A) $K_a \times K_b = 1$ (B) $K_a \times K_b = 10^{-14}$

7. In a 0.1 M solution of H_2CO_3 , $[\text{H}_2\text{CO}_3] = 0.09 \text{ M}$, $[\text{HCO}_3^-] = 0.01 \text{ M}$, $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$. What is the K_a of H_2CO_3 ?
- (A) 1.1×10^{-4} (B) 1.1×10^{-5} (C) 1.1×10^{-6} (D) 1.1×10^{-7}

8. A solution of acetic acid and its conjugate base, sodium acetate, is prepared with $[\text{CH}_3\text{COOH}] = 0.1 \text{ M}$ and $[\text{CH}_3\text{COO}^-] = 0.1 \text{ M}$. What is the pH of this solution? (correct)
- (A) 4 (B) 5 (C) 6 (D) 7



9. In a 0.1 M solution of H_2CO_3 , $[\text{H}_2\text{CO}_3] = 0.09 \text{ M}$, $[\text{HCO}_3^-] = 0.01 \text{ M}$, $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$. What is the K_a of H_2CO_3 ?
- (A) 1.1×10^{-4} (B) 1.1×10^{-5} (C) 1.1×10^{-6} (D) 1.1×10^{-7}

10. In a 0.1 M solution of acetic acid, CH_3COOH , $[\text{CH}_3\text{COOH}] = 0.09 \text{ M}$, $[\text{CH}_3\text{COO}^-] = 0.01 \text{ M}$, $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$. What is the K_a of CH_3COOH ?
- (A) 1.1×10^{-4} (B) 1.1×10^{-5} (C) 1.1×10^{-6} (D) 1.1×10^{-7}



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M Carnoy



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